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## **Redox Practice Problems With Answers**

Practice: Redox reactions questions. This is the currently selected item.

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Oxidizing and reducing agents.

Disproportionation.

Worked example:

Balancing a redox equation in acidic solution. Worked

example: Balancing a redox equation in basic solution.

## **Redox reactions questions (practice)**

**| Khan Academy**

Practice Problems:

Redox Reactions

(Answer Key)

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Determine the oxidation number of the elements in each of the following compounds: a.  $\text{H}_2\text{CO}_3$   
H: +1, O: -2, C: +4

## **Practice Problems: Redox Reactions (Answer Key)**

Problem #8:  $\text{Fe} + \text{HCl} \rightarrow \text{HFeCl}_4 + \text{H}_2$ .

Solution: 1) This problem poses interesting problems, especially with the Cl. The key to solving this

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problem is to eliminate everything not directly involved in the redox. That means the H in  $\text{HFeCl}_4$  as well as the Cl in it and HCl. When we do that, this is the unbalanced, ionic form we wind up with:

## **Balancing redox reactions in acidic solution: Problems #1-10**

Base your answer(s) to the following question(s) on the

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diagram below, which represents a voltaic ...  
page 7 Redox practice worksheet. Problem-  
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## **Redox practice worksheet**

Practice Problems:  
Redox Reactions.  
Determine the  
oxidation number of  
the elements in each of



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the following  
compounds: a.  $\text{H}_2\text{CO}_3$  b.  $\text{N}_2$  c.  $\text{Zn(OH)}_2$  d.  $\text{NO}_2$  e.  $\text{LiH}$  f.  $\text{Fe}_3\text{O}_4$   
Hint; Identify the species being oxidized and reduced in each of the following reactions:  
a.  $\text{Cr} + \text{Sn}^{4+} \rightarrow \text{Cr}^{3+} + \text{Sn}^{2+}$  b.  $3 \text{Hg}^{2+} + 2 \text{Fe (s)} \rightarrow 3 \text{Hg} + 2 \text{Fe}^{3+}$  c.  $2 \text{As} \dots$

## **Practice Problems: Redox Reactions**

to, the revelation as capably as perspicacity

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of this redox reactions  
problems with answers  
can be taken as  
competently as picked  
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Oxidation-Reduction  
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Reaction Problems  
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Base and ...

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## **Redox Reactions Problems With Answers | happyhounds ...**

Practice Problems;  
References; Oxidation-  
Reduction or "redox"  
reactions occur when  
elements in a chemical  
reaction gain or lose  
electrons, causing an  
increase or decrease in  
oxidation numbers. The  
Half Equation Method  
is used to balance  
these reactions. In a  
redox reaction, one or

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more element becomes oxidized, and one or more element becomes ...

## **Balancing Redox Reactions: Examples - Chemistry LibreTexts**

When balancing redox reactions, the overall electronic charge must be balanced in addition to the usual molar ratios of the component reactants and products. This

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example problem illustrates how to use the half-reaction method to balance a redox reaction in a solution.

## **Balance Redox Reaction Example Problem**

Interactive Practice Problems: Questions are randomly generated from the question sets written by Angela King. You can continue doing the

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Practice Problems until you have got all the available questions correct. You get points for correct answers. If you get it wrong, the question may appear again if you continue.

## **Dr. Angela King's Online Redox Tutorial and Practice Problems**

We know about redox reactions, and we even learned how to balance them under both acidic

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and basic conditions.

Time for a quiz!

Balance this redox  
reaction. ...

## **Practice Problem: Balancing Redox Reactions - YouTube**

Redox Balancing  
Practice. The following  
are a series of fill-in  
reviews for balancing  
redox problems. Two of  
them focus on the step-  
by-step methods for  
balancing, while the  
others require only the

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overall balanced equation. You can do them individually, or start anywhere in the sequence and move forward, backward, or back to this page.

## **Redox Balancing Practice -**

### **ScienceGeek.net**

Problem #5: A rock sample is to be assayed for its tin content by an oxidation-reduction titration with  $\text{I}_3^- (\text{aq})$ . A 10.00 g



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A sample of the rock is crushed, dissolved in sulfuric acid, and passed over a reducing agent so that all the tin is in the form  $\text{Sn}^{2+}$ . The  $\text{Sn}^{2+}$  (aq) is completely oxidized by 34.60 mL of a 0.5560 M solution of  $\text{NaIO}_3$ . The balanced equation for the reaction is

## **ChemTeam: Redox Titration**

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questions. This is the  
currently selected  
item. Acid-base  
titrations. Worked  
example: Determining  
solute concentration by  
acid-base titration.

## **Titration questions (practice) | Titrations | Khan Academy**

Practice Problems  
Electrochemistry. 1.  
What is the difference

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between an oxidation-reduction reaction and a half-reaction? 2.

What is the function of the salt bridge in an electrochemical cell? 3.

What is the criterion for spontaneous chemical change based on cell potentials? Explain. 4.

## **CHM 112 Electrochemistry Practice Problems**

Chapter 20 Worksheet:  
Redox I. Determine  
what is oxidized and

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what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

- $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$
- $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$
- $2\text{Cs} + \text{Br}_2 \rightarrow 2\text{CsBr}$
- $3\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$
- $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
- $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- $\text{Si} + 2\text{F}_2 \rightarrow \text{SiF}_4$
- $2\text{Ca} + \text{O}_2 \rightarrow 2\text{CaO}$
- 

## Chapter 20

### Worksheet Redox

Redox Unit Test

Practice Problems ...

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Redox Practice

Problem Answers . 1.

Consult book and  
notes. 2.  $3\text{H}_2\text{S} +$

$2\text{HNO}_3 \rightarrow 3\text{S} + 2\text{NO} +$

$4\text{H}_2\text{O}$  +1 -2 +1+5-2 0

+2-2 +1-2 . Sulfur is

being oxidized and the  
reducing agent .

Nitrogen is being

reduced and the

oxidizing agent . 3.  $3\text{S}$

...

**Redox Practice  
Problems - Weebly**

9:29 Balancing Redox

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Reactions Practice  
Worksheet Key

Problem #3. Balancing  
a redox reaction under  
basic conditions.

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Repatriation in the  
1930's. Use this to  
answer the warm-up  
from Wednesday 1/13  
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