

## Pearson Education Chapter 8 Covalent Bonding Answers

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Pearson Education Chapter 8 Covalent Chapter 8 Covalent Bonding 183 Section Review Objectives • State a rule that usually tells how many electrons are shared to form a covalent bond • Describe how electron dot formulas are used • Predict when two atoms are likely to be joined by a double or a triple covalent bond • Distinguish

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8.2 The Nature of Covalent Bonding > 23 Copyright © Pearson Education, Inc., or its affiliates. All Rights Reserved. • Experimental evidence, however, indicates ...

### Chapter 8

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sharing a pair of electrons betw.... The smallest particle of a substance that retains the chemical.... A molecule that consists of two atoms of the same element that....

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Exercise 2 Lewis Structure of a Compound. Solve. Nitrogen must share a pair of electrons with three fluorine atoms to complete its octet. Thus, the binary compound these two elements form must be NF<sub>3</sub>. Given the Lewis symbols for nitrogen and fluorine in Table 8.1, predict the formula of the stable binary compound (a compound composed of two elements) formed when nitrogen reacts with ...

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